

REMARKS

The applicants thank the Examiner for the thorough examination of the application. Attached is a copy of D. Owiny et al., J. Organometallic Chemistry 678 (2003) 134-141. No new matter is believed to be added to the application by this amendment.

Summary of Interview

The Examiner is thanked for graciously conducting an interview with the Applicants' representative on August 3, 2004. During the Interview, enablement and repetitive claim language in claims 1-14 were discussed. The discussion included observations that enablement for binding of only one bidentate ligand may be shown by a prior art literature synthesis.

The Interview Summary prepared by the Examiner has been reviewed, and it appears to accurately reflect the substance of the interview.

Status of the Claims

Claims 1-15 are pending in the application. Claim 15 is withdrawn from consideration by the Examiner. The claims have been amended to remove redundant language. The scope of the claims has not been narrowed by these amendments.

Rejection Under 35 U.S.C. §112, First Paragraph (Paragraph 3 of the Office Action)

Claims 1-14 are rejected under 35 U.S.C. §112, first paragraph, as failing to comply with the written description requirement. Applicants respectfully traverse.

At page 2 of the Office Action, the Examiner asserts: "The claims as presently drawn only permit one of the bidentate ligands be bonded to the transition metal. However, throughout the entire specification there is not a single working prophetic example of a transition metal compound having only one bidentate ligand bonded to it." The Examiner then takes the position: "This amounts to a failure to teach how to make the claimed invention, instead being a mere invitation to experiment how to make the invention as claimed."

However, the single bidentate ligand binding embodied in the invention can be achieved without undue experimentation by one having ordinary skill.

"The test of enablement is whether one reasonably skilled in the art could make or use the invention from the disclosures in the patent coupled with information known in the art without undue experimentation." United States v. Teletronics, Inc., 8 USPQ2d 1217 (Fed. Cir. 1988); In re Stephens, 188 USPQ 659 (CCPA 1976). "A patent may be enabling even though some experimentation is necessary; the amount of experimentation, however, must not be unduly extensive." Utter v. Hiraga, 6 USPQ2d 1709, 1714 (Fed. Cir. 1988).

As evidence of the general knowledge of the synthesis of a complex of a metal with a single bidentate ligand, please find attached a copy of D. Owiny et al., J. Organometallic Chemistry 678 (2003) 134-141 entitled *Synthesis, Structural*

Determination, and Ethylene Polymerization Chemistry of Mono(salicylaldiminato) Complexes of Titanium (IV). This article describes the synthesis of a compound having a single bidentate ligand bonded to the transition metal. Although the article was published in 2003, references [3] and [4] of the article contain references to publications as early as 1991 that pertain to mono(salicylaldiminato)titanium (IV). A person having ordinary skill in the art would be cognizant of these references and would hence be able to bind a single bidentate ligand to a transition metal without undue experimentation.

This reference by D. Owiny et al. also satisfies the Examiners comments in the Interview Summary of August 3, 2004, which states: "Enablement for compounds having only one bidentate ligand may be shown via prior art literature synthesis. This would show how one of ordinary skill in the art would be able to practice the invention without undue experimentation."

As a result, the claims are in full compliance with the written description requirement of 35 U.S.C. §112, first paragraph. This rejection is overcome and withdrawal thereof is respectfully requested.

Rejection Under 35 U.S.C. §112, Second Paragraph (Paragraph 4 of the Office Action)

Claims 3-8 are rejected under 35 U.S.C. §112, second paragraph as being indefinite. Applicants respectfully traverse.

In the Office Action, the Examiner notes prolix and repetitive subject matter. The Examiner's comments have been considered. The claims have been amended

to remove superfluous formulas and recitations. As a result, the amended claims are clear, definite and have full antecedent basis.

This rejection is overcome and withdrawal thereof is respectfully requested.

The Drawings

The Examiner has accepted the drawing figures in the Office Action mailed June 22, 2004.

Information Disclosure Statements

The Applicants thank the Examiner for considering the Information Disclosure Statements filed August 31, 2001 and November 3, 2003 and for making the initialed PTO-1449 forms of record in the application in the Office Action mailed February 9, 2004.

Prior Art Cited by the Examiner

The prior art previously cited but not utilized by the Examiner indicates the status of the conventional art that the invention supercedes. Additional remarks are accordingly not necessary.

Foreign Priority

The Examiner has acknowledged the claim for foreign priority.


Conclusion

Should there be any outstanding matters that need to be resolved in the present application, the Examiner is respectfully requested to contact Robert E. Goozner, Ph.D. (Reg. No. 42,593) at the telephone number of the undersigned below, to conduct an interview in an effort to expedite prosecution in connection with the present application.

If necessary, the Commissioner is hereby authorized in this, concurrent, and future replies, to charge payment or credit any overpayment to Deposit Account No. 02-2448 for any additional fees required under 37 C.F.R. §§ 1.16 or 1.17; particularly, extension of time fees.

Respectfully submitted,

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Attachment(s): D. Owiny et al., J. Organometallic Chemistry 678 (2003) 134-141



Synthesis, structural determination, and ethylene polymerization chemistry of mono(salicylaldiminato) complexes of titanium(IV)

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Abstract

Titanium(IV)mono(salicylaldiminato) complexes $[L^1Ti(NMe_2)_3]$ (1) and $[L^1TiCl_3]$ (2) have been synthesized by treatment of $Ti(NMe_2)_4$ or $TiCl_4$ with one equivalent of [4, 6-Bu₂-2-(CH=NBu')C₆H₃OH] (**L¹H**) or [4, 6-Bu₂-2-(CH=NBu')C₆H₃OSiMe₃] (**L¹SiMe₃**), respectively. The compounds are monomeric in solution and in the solid-state. Reactions of $TiCl_4$ with one equivalent of [4, 6-Bu₂-2-(CH=NCH₂Ph)C₆H₃OH] (**L²H**) and [4, 6-Bu₂-2-(CH=N(2-C₆H₄OH))C₆H₃OH] (**L³H₂**) produced $[L^2TiCl_2(\mu-Cl)]_2$ (3) and $[L^3TiCl_2]_2$ (4), respectively. $[L^3TiCl_2(THF)]$ (5) was also produced in quantitative yield when 4 was stirred in THF for 16 h. The reaction of $TiCl_4$ with **L³H₂** (Two equivalents) in toluene gave $[(L^3)_2Ti]$ (6). The molecular structures of 2–4 and 6 were established by single-crystal X-ray diffraction studies; and 4 displayed a rare face to face π – π stacking interaction in its structure. Compounds 1 and 2 showed modest ethylene polymerization activities at 25 °C with 900 molar equivalents of methylalumoxane (MAO) as co-catalyst.

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Keywords: Titanium(IV) complexes; Olefin polymerization; Salicylaldiminato ligands; Schiff base ligands

1. Introduction

There has been a great deal of interest in the development of well-defined transition metal catalysts for the polymerization of α -olefins since the discovery of highly active Group 4 metallocene catalysts [1]. A wide variety of ligand environments and transition metals have been investigated and olefin polymerization catalysts based on both early- and late transition metals have been developed, some of which show activities superior or comparable to those of Group 4 metallocenes [2]. In fact, Group 4 metal complexes based on chelating di(amido)- 2c2d2e2f, amine-bis(phenolato)- 2n, or bis(salicylaldiminato) 2h2i2j2k ligands have furnished highly active and/or living olefin polymerization catalysts. However, despite extensive investigation of organometallic complexes of salicylaldiminato ligands and the utility of bis(salicylaldiminato)titanium(IV) com-

plexes in olefin polymerization chemistry, few mono(salicylaldiminato)titanium(IV) complexes have been described and their reaction chemistry is poorly developed [3,4]. This is surprising since salicylaldimines (HOC₆H₄CH=NR, R = alkyl or aryl) are obtained by easy synthetic routes and hence their steric and electronic properties can be readily tuned, via substitution of the aromatic ring or modification of the imino nitrogen substituent 2h2d2e2f2g2h2k[3]. Herein, we describe the synthesis, X-ray crystallographic characterization, and ethylene polymerization chemistry of mono(salicylaldiminato)titanium(IV) complexes $[\{ 4, 6-Bu'_2-2-(CH=NBu')C_6H_3O \} TiCl_3]$ (1) and $[\{ 4, 6-Bu'_2-2-(CH=NBu')C_6H_3O \} Ti(NMe_2)_3]$ (2).

2. Experimental

2.1. General

$Ti(NMe_2)_4$ [5], $TiCl_4(THF)_2$ [6], and salicylaldimine compounds, 4,6-bis(*tert*-butyl)-2-((*tert*-butyl)imino-methyl)phenol [4, 6-Bu'₂-2-(CH=NBu')C₆H₃OH] (**L¹H**)

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[7], 4,6-bis(*tert*-butyl)-2-[(benzyl)iminomethyl]phenol [4,6-Bu₂-2-(CH=NCH₂Ph)C₆H₃OH] (**L²H**) [8], and 4,6-bis(*tert*-butyl)-2-[(2-hydroxyphenyl)imino-methyl]phenol [4,6-Bu₂-2-(CH=N(2-C₆H₄OH))C₆H₃OH] (**L³H₂**) [9], were prepared via modifications of literature methods. All other experiments were performed under dry nitrogen atmosphere using standard Schlenk techniques or in a MBraun glovebox. Benzene-*d*₆, toluene, tetrahydrofuran, pentane, and heptane were distilled from sodium benzophenone ketyl (with 1 ml l⁻¹ of triethyleneglycol dimethyl ether added as a solubilizing agent in the case of pentane and heptane). CD₂Cl₂ and CDCl₃ were distilled from calcium hydride. All solvents were stored in the glovebox over 4A molecular sieves that had been dried under vacuum at 150 °C for at least 48 h prior to use. All other chemicals were purchased from Aldrich Chemical Co. and used without further purification (unless otherwise stated). ¹H- and ¹³C-NMR spectra were recorded on a Varian Gemini-200 spectrometer or a Varian VXR-400 spectrometer at ca. 22 °C. ¹H- and ¹³C-chemical shifts were referenced to residual solvent peaks. Infrared spectra were recorded on a Nicolet Magna 560 spectrometer. Mass spectral data were obtained from the University of Kentucky Mass Spectrometry Center on a Thermo Finnigan (San Jose, CA) Polaris Q (quadrupole ion trap) spectrometer. Elemental analyses were performed by Complete Analysis Laboratories, Inc., Parsippany, NJ.

2.2. Synthesis of [4,6-Bu₂-2-(CH=NBu^t)C₆H₃OSiMe₃] (**L¹SiMe₃**)

A Et₂O (5 mL) solution of KOBu^t (257 mg, 2.29 mmol) was added in two portions to a Et₂O (10 mL) solution of 4,6-bis(*tert*-butyl)-2-[(*tert*-butyl)iminomethyl]phenol (**L¹H**, 54.4 mg, 1.88 mmol) at room temperature (r.t.). The reaction mixture was allowed to stir for 0.5 h, after which the yellowish-white precipitate was collected by filtration under N₂ atmosphere. After washing the precipitate several times with Et₂O, it was put back into Et₂O. Me₃SiCl (0.25 ml, 1.97 mmol) was added to the Et₂O suspension and the resulting mixture was stirred for 20 min. The reaction mixture was stripped to dryness under reduced pressure, the residue was extracted with pentane, and the solvent was removed under vacuum to give a yellowish-white crystalline solid. Yield: 0.510 g, 75%. ¹H-NMR (CDCl₃): δ 8.51 (s, 1H, Bu^tN=CH), 7.73 (d, 1H, *J*_{HH} = 6.0 Hz, arom CH), 7.41 (d, 1H, *J*_{HH} = 5.6 Hz, arom CH), 1.41 (s, 9H, Bu^t), 1.32 (s, 18H, Bu^t, NBu^t), 0.32 (s, 9H, SiMe₃). ¹³C-NMR (CDCl₃): δ 153.8 (Bu^tN=CH), 152.3, 143.7, 140.2, 128.8, 127.1, 123.1, 57.8 (NCMe₃), 34.6 (CMe₃), 34.3 (CMe₃), 31.6 (CMe₃), 30.9 (CMe₃), 30.0 (NCMe₃), 2.3 (SiMe₃).

2.3. Synthesis of titanium compounds 1–6

2.3.1. [**L¹Ti(NMe₂)₃**] (**1**)

A toluene (15 mL) solution of **L¹H** (0.381 g, 1.32 mmol) was added drop-wise to a stirred toluene (8 mL) solution of [Ti(NMe₂)₄] (0.297 g, 1.33 mmol) at r.t. After completion of the addition, the yellow–orange reaction mixture was stirred for 20 min. The solution was stripped under reduced pressure to give a yellow–orange crystalline solid. Yield: 0.538 g, 92%. ¹H-NMR (C₆D₆): δ 8.52 (s, 1H, Bu^tN=CH), 7.73 (d, 1H, *J*_{HH} = 2.4, arom CH), 7.33 (d, 1H, *J*_{HH} = 2.4, arom CH), 3.20 (s, 18H, NMe₂), 1.66 (s, 9H, Bu^t), 1.38 (s, 9H, Bu^t), 1.06 (s, 9H, NBu^t). ¹³C-NMR (C₆D₆): δ 166.5 (Bu^tN=CH), 160.9, 140.2, 139.7, 129.3, 128.7, 122.6, 60.6 (NCMe₃), 46.7 (NMe₂), 35.9 (CMe₃), 34.6 (CMe₃), 32.0 (CMe₃), 31.2 (CMe₃), 30.1 (NCMe₃). IR (CH₂Cl₂, cm⁻¹): 1612 ν(C=N). Anal. Calc. for C₂₅H₄₈N₄O₂Ti: C, 64.09; H, 10.32; N, 11.96. Found: C, 63.97; H, 10.10; N, 11.67%.

2.3.2. [**L¹TiCl₃**] (**2**)

A toluene (12 mL) solution of **L¹SiMe₃** (1.07 g, 3.27 mmol) was added drop-wise to a stirred toluene (20 mL) solution of TiCl₄ (0.830 g, 4.36 mmol) at 0 °C (ice bath). Upon complete addition, the solution was stirred (0–2 °C) for 4 h. The resulting orange–brown precipitate was collected by filtration at r.t., washed once with toluene (10 mL), and dried under vacuum. Yield: 0.830 g, 58%. ¹H-NMR (CDCl₃): δ 8.59 (s, 1H, Bu^tN=CH), 7.66 (d, 1H, *J*_{HH} = 2.4, arom CH), 7.44 (d, 1H, *J*_{HH} = 2.4, arom CH), 1.76 (s, 9H, Bu^t), 1.53 (s, 9H, Bu^t), 1.37 (s, 9H, NBu^t). ¹³C-NMR (CDCl₃): δ 163.7 (Bu^tN=CH), 157.9, 144.7, 141.6, 134.1, 127.2, 118.2, 60.9 (NCMe₃), 35.4 (CMe₃), 35.0 (CMe₃), 31.5 (CMe₃), 30.1 (CMe₃), 28.3 (NCMe₃). IR (CH₂Cl₂, cm⁻¹): 1604 ν(C=N). Anal. Calc. for C₁₉H₃₀Cl₃NOTi: C, 51.55; H, 6.83. Found: C, 51.52; H, 6.82%.

2.3.3. [**L²TiCl₂(μ-Cl)**]₂ (**3**)

A heptane (18 mL) solution of **L²H** (1.15 g, 3.55 mmol) was added drop-wise to a stirred heptane (30 mL) solution of TiCl₄ (0.690 g, 3.64 mmol) at –78 °C. After the addition was complete, the brick-red reaction mixture was allowed to warm gradually up to r.t. and stirred for ~10 h. The resulting red–brown precipitate was collected by filtration, washed with pentane (4 × 15 mL), and dried under vacuum to give an orange–red powder. Yield: 3.01 g, 89%. ¹H-NMR (CDCl₃): δ 8.22 (s, 1H, PhCH₂N=CH), 7.65 (d, 1H, *J*_{HH} = 2.0, arom CH), 7.55–7.34 (m, 5H, PhCH₂), 7.16 (d, 1H, *J*_{HH} = 2.0, arom CH), 5.36 (s, 2H, PhCH₂), 1.54 (s, 9H, Bu^t), 1.30 (s, 9H, Bu^t). ¹³C-NMR (CDCl₃): δ 166.1 (PhCH₂N=CH), 161.6, 148.4, 136.8, 135.6, 131.3, 130.6, 129.5, 129.1, 129.0, 124.9, 63.2 (PhCH₂), 35.4 (CMe₃), 35.1 (CMe₃), 31.4 (CMe₃), 29.76 (CMe₃). MS (EI, 70 eV, *m/z*): 869 [M⁺–Cl]. IR (CHCl₂, cm⁻¹):

1613 $\nu(\text{C}=\text{N})$. Anal. Calc. for $\text{C}_{44}\text{H}_{56}\text{Cl}_6\text{N}_2\text{O}_2\text{Ti}_2$: C, 55.43; H, 5.92. Found: C, 55.26; H, 5.73%.

2.3.4. $[\text{L}^3\text{TiCl}_2]_2$ (4)

A toluene (20 ml) solution of L^3H_2 (0.435 g, 1.33 mmol) was added drop-wise to a stirred toluene (5 ml) solution of TiCl_4 (0.259 g, 1.36 mmol) at -78°C . After the addition was complete, the initially orange–red reaction mixture was allowed to warm gradually up to r.t. and stirred for ~ 10 h. The resulting brown mixture was filtered and the precipitate was washed with pentane (3×15 ml), then dried under vacuum. Yield: 0.923 g, 79%. ^1H -NMR (CDCl_3): δ 8.82 (s, 2H, $\text{N}=\text{CH}$), 7.72 (d, 2H, $J_{\text{HH}} = 2.4$, arom CH), 7.38 (d, 2H, $J_{\text{HH}} = 2.2$ arom CH), 7.34 (d, 2H, $J_{\text{HH}} = 8.0$, arom CH), 7.26 (t, 2H, $J_{\text{HH}} = 8.0$ arom CH), 6.95 (t, 2H, $J_{\text{HH}} = 8.0$, arom CH), 6.82 (d, 2H, $J_{\text{HH}} = 8.0$, arom CH), 1.55 (s, 18H, Bu'), 1.37 (s, 18H, Bu'). ^{13}C -NMR (CDCl_3): δ 157.9 ($\text{N}=\text{CH}$), 146.7, 137.3, 133.6, 131.2, 129.5, 129.3, 128.4, 125.5, 122.5, 115.1, 114.7, 35.5 (CMe_3), 34.86 (CMe_3), 31.39 (CMe_3), 29.92 (CMe_3). MS (EI, 70 eV, m/z): 849 [$\text{M}^+ - \text{Cl}$]. IR (CH_2Cl_2 , cm^{-1}): 1600 $\nu(\text{C}=\text{N})$. A sample of 4 was recrystallized from CHCl_3 prior to microanalysis and gave $[\text{L}^3\text{TiCl}_2]_2 \cdot 2\text{CHCl}_3$. Anal. Calc. for $\text{C}_{44}\text{H}_{52}\text{Cl}_{10}\text{N}_2\text{O}_4\text{Ti}_2$: C, 47.07; H, 4.64; N, 2.49. Found: C, 47.64; H, 4.99; N, 2.23%.

2.3.5. $[\text{L}^3\text{TiCl}_2(\text{THF})]$ (5)

A toluene (6 ml) solution of L^3H_2 (0.155 g, 0.476 mmol) was added drop-wise to a stirred toluene (5 ml) suspension of $[\text{TiCl}_4(\text{THF})_2]$ (0.153 g, 0.459 mmol) at -78°C . After the addition was complete, the initially orange–red reaction mixture was allowed to warm gradually up to r.t. and stirred for ~ 16 h. The resulting brown–black mixture was filtered and the precipitate was washed with pentane (2×10 ml), then dried under vacuum. Yield: 0.137 g, 58%. ^1H -NMR (CDCl_3): δ 8.46 (s, 1H, $\text{N}=\text{CH}$), 7.55 (d, 1H, $J_{\text{HH}} = 2.2$, arom CH), 7.26 (t, 1H, $J_{\text{HH}} = 8.0$, arom CH), 7.18 (d, 1H, $J_{\text{HH}} = 2.0$, arom CH), 7.09 (t, 1H, $J_{\text{HH}} = 8.0$, arom CH), 6.80 (t, 1H, $J_{\text{HH}} = 8.0$, arom CH), 6.49 (d, 1H, $J_{\text{HH}} = 8.0$, arom CH), 4.20 (m, 4H, THF), 1.93 (m, 4H, THF), 1.47 (s, 9H, Bu'), 1.33 (s, 9H, Bu'). ^{13}C -NMR (CDCl_3): δ 162.0 ($\text{N}=\text{CH}$), 160.6, 156.1, 145.4, 139.0, 136.7, 133.0, 131.2, 129.7, 124.2, 122.0, 114.6, 114.2, 73.2 (THF), 35.3 (CMe_3), 34.7 (CMe_3), 31.4 (CMe_3), 30.0 (CMe_3), 25.6 (THF). IR (CH_2Cl_2 , cm^{-1}): 1604 $\nu(\text{C}=\text{N})$. Anal. Calc. for $\text{C}_{25}\text{H}_{33}\text{Cl}_2\text{NO}_3\text{Ti}$: C, 58.39; H, 6.47; N, 2.72. Found: C, 58.34; H, 6.22; N, 2.92%.

2.3.6. $[(\text{L}^3)_2\text{Ti}]$ (6)

A toluene (6 ml) solution of L^3H_2 (0.256 g, 0.786 mmol) was added drop-wise to a stirred toluene (6 ml) solution of TiCl_4 (77.0 mg, 0.409 mmol) at -78°C . After the addition was complete, the initially red reaction mixture was allowed to warm gradually up to

r.t. and stirred for ~ 16 h. The resulting red mixture was filtered and the precipitate was washed with pentane (3×15 ml), then dried under vacuum. Yield: 0.202 g, 71%. ^1H -NMR (CDCl_3): δ 8.86 (s, 2H, $\text{N}=\text{CH}$), 7.46 (s, 2H, arom CH), 7.34 (d, 2H, $J_{\text{HH}} = 8.0$, arom CH), 7.32 (s, 2H, arom CH), 7.04 (t, 2H, $J_{\text{HH}} = 8.0$, arom CH), 6.75 (t, 2H, $J_{\text{HH}} = 8.0$, arom CH), 6.56 (d, 2H, $J_{\text{HH}} = 8.0$, arom CH), 1.34 (s, 18H, Bu'), 1.13 (s, 18H, Bu'). ^{13}C -NMR (CDCl_3): δ 161.5 ($\text{N}=\text{CH}$), 158.0, 142.2, 138.7, 136.6, 131.9, 130.1, 128.4, 121.7, 119.4, 115.2, 114.2, 35.0 (CMe_3), 34.5 (CMe_3), 31.5 (CMe_3), 29.4 (CMe_3). IR (CH_2Cl_2 , cm^{-1}): 1604 $\nu(\text{C}=\text{N})$. MS (EI, 70 eV, m/z): 694 [M^+]. A sample of 6 was recrystallized from CHCl_3 prior to microanalysis. The resulting crystals contained lattice CHCl_3 molecules, \sim one-third of a CHCl_3 molecule per titanium. Anal. Calc. for $\text{C}_{42}\text{H}_{50}\text{N}_2\text{O}_4\text{Ti} \cdot (\text{CHCl}_3)_{0.33}$: C, 69.20; H, 6.90; N, 3.81. Found: C, 69.58; H, 6.96; N, 4.03%.

2.4. Ethylene polymerization studies

Toluene (60 ml) was charged into a 250 ml two-necked Schlenk flask under N_2 atmosphere in a glove-box. An excess of MAO (Al/Ti ratio = 900/1) was added and the solution was stirred for several minutes. Next, the catalyst precursor (0.027 mmol of Ti) was added as a solid and the solution was stirred for 30 min. The N_2 atmosphere was replaced with ethylene (passed through a column of BASF catalyst R3-11 and a solution of MAO) and the pressure was maintained at 1 atmosphere by slow bubbling through the solution and controlling the pressure with a bubbler at the outlet. The polymerization was carried out for 10 min. Polymerizations were quenched with MeOH (30 ml) and then 1 M HCl solution (30 ml). The resulting suspension was vigorously stirred until both layers were colorless and clearly separated (~ 10 min). Polyethylene was filtered off, washed with 1 M HCl, and MeOH then dried at 70°C for 48 h.

2.5. Crystallographic study

The crystal data for $[\text{L}^2\text{TiCl}_2(\mu\text{-Cl})]_2$ (3), $[\text{L}^3\text{TiCl}_2]_2$ (4), and $[(\text{L}^3)_2\text{Ti}]$ (6) are collected in Table 1. Further details of the crystallographic study are given in Section 5.

3. Results and discussion

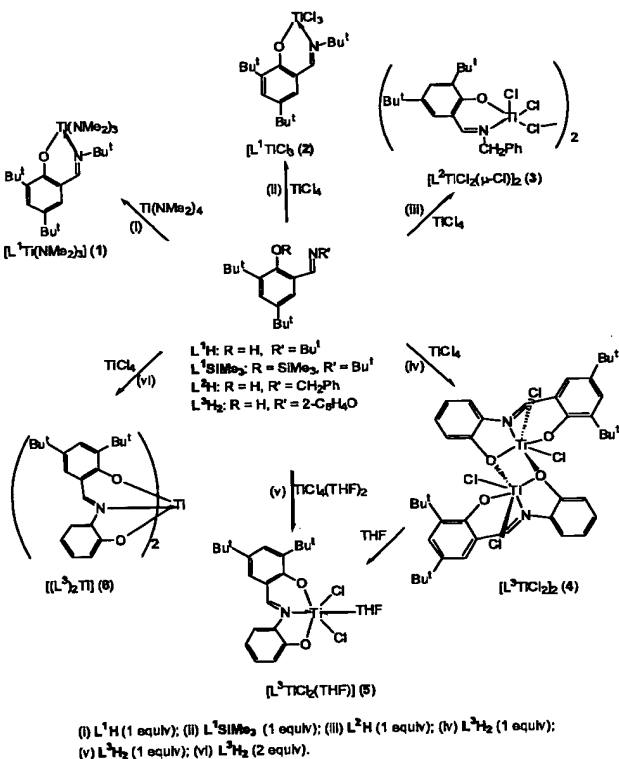
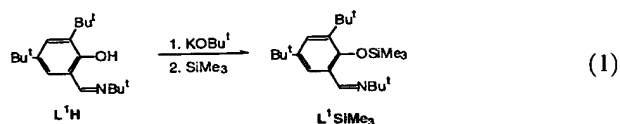
The synthesis of the salicylaldiminato complexes of titanium described in this report is summarized in Scheme 1. The mono(salicylaldiminato)titanium(IV) complex $[\text{L}^1\text{Ti}(\text{NMe}_2)_3]$ (1) was isolated in excellent yield from reaction of $\text{Ti}(\text{NMe}_2)_4$ with one equivalent of $[4,6\text{-Bu}'_2\text{-2-(CH}=\text{N}^+\text{Bu}')\text{C}_6\text{H}_3\text{OH}]$ (L^1H) in toluene at \sim

Table 1
Crystallographic data for 3·C₅H₁₂, 4·CHCl₃, and 6·CH₂Cl₂

	3·C ₅ H ₁₂	4·CHCl ₃	6·CH ₂ Cl ₂
Empirical formula	C ₄₉ H ₆₈ Cl ₆ N ₂ O ₂ Ti ₂	C _{42.50} H _{50.50} Cl _{5.50} N ₂ O ₄ Ti ₂	C ₄₃ H ₅₂ Cl ₂ N ₂ O ₄ Ti
Formula weight	1025.55	944.12	779.67
Temperature (K)	150(2)	90(2)	90(2)
Crystal system	Monoclinic	Tetragonal	Triclinic
Space group	<i>P</i> 2 ₁ / <i>c</i>	<i>P</i> -421 <i>c</i>	<i>P</i> $\bar{1}$
<i>a</i> (Å)	12.4720(5)	15.9770(2)	11.7557(6)
<i>b</i> (Å)	21.1090(8)	15.9770(2)	14.2041(7)
<i>c</i> (Å)	19.7410(7)	17.5674(3)	14.2328(7)
α (°)	90	90	104.837(3)
β (°)	90.6330(18)	90	113.565(3)
γ (°)	90	90	99.152(3)
<i>V</i> (Å ³)	5196.9(3)	4484.33(11)	2012.01(17)
<i>Z</i>	4	4	2
<i>D</i> _{calc} (g cm ⁻³)	1.311	1.398	1.287
Final <i>R</i> indices [<i>I</i> > 2σ(<i>I</i>): <i>R</i> ₁ , <i>wR</i> ₂	0.0789, 0.1598	0.0378, 0.0717	0.0892, 0.1456
<i>wR</i> ₂ , <i>R</i> ₁ (all data)	0.1033, 0.1695	0.0470, 0.0741	0.1207, 0.1544

25 °C. The reaction between TiCl₄ and L¹H (one equivalent) in heptane did not cleanly produce mono(salicylaldiminato)titanium trichloride [L¹TiCl₃] (2) [10]. Instead, 2 was obtained in excellent yield from the reaction of TiCl₄ with one equivalent of [4,6-Bu^t₂-2-(CH=NBu^t)C₆H₃OSiMe₃] (L¹SiMe₃) in toluene at 0 °C. Clean silylation of L¹H was achieved as

shown in Eq. (1). Reaction of TiCl₄ with one equivalent of [4,6-Bu^t₂-2-(CH=NCH₂Ph)C₆H₃OH] (L²H) or [4,6-Bu^t₂-2-(CH=N(2-C₆H₄OH))C₆H₃OH] (L³H₂) at -78–25 °C produced [L²TiCl₂(μ-Cl)]₂ (3) and [L³TiCl₂]₂ (4), respectively. The fact that 3 and 4 are dimers likely reflects the reduced steric constraint on titanium by the respective salicylaldimine ligand in comparison to [4,6-Bu^t₂-2-(CH=NBu^t)C₆H₃O⁻] (L¹). Consistent with this suggestion, the reaction of TiCl₄(THF)₂ with L³H₂ (one equivalent) in toluene afforded monomeric [L³TiCl₂(THF)] (5), which was also produced in quantitative yield when 4 was dissolved in THF and stirred for 16 h. [(L³)₂Ti] (6) was obtained in good yield from the reaction between TiCl₄ and L³H₂ (two equivalents) in toluene at -78–25 °C.



Scheme 1.

All of the compounds 1–6 are air- and moisture-sensitive, thermally stable red to red-brown solids. They are conveniently stored in the solid-state under N₂ atmosphere at ambient temperature without any observable decomposition. All of the compounds are readily soluble in polar hydrocarbon solvents, such as THF, chloroform, and dichloromethane, and are practically insoluble in aliphatic hydrocarbon solvents, such as pentane and heptane. While 1 is quite soluble in aromatic hydrocarbon solvents, such as benzene and toluene, 2–6 are only sparingly soluble in these solvents. The formulations proposed for 1–6 were confirmed by microanalysis, ¹H- and ¹³C-NMR, and/or mass spectrometry (see Section 2). The molecular structures of [L¹TiCl₃] (2), [L²TiCl₂(μ-Cl)]₂ (3), [L³TiCl₂]₂ (4), and

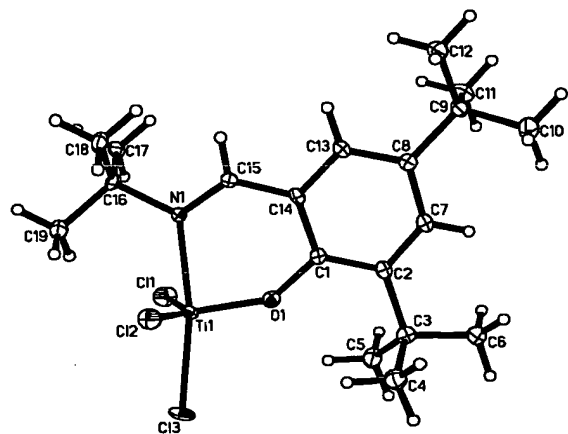


Fig. 1. An ORTEP diagram of the molecular structure of **2** showing 50% thermal ellipsoid probabilities.

$[(L^3)_2Ti]$ (**6**) were also established by single-crystal X-ray diffraction studies (Figs. 1–4). Selected metrical parameters are listed in Tables 2 and 3. Room temperature 1H - and ^{13}C -NMR data for $[L^1Ti(NMe_2)_3]$ (**1**) indicate fast exchange of the NMe_2 ligands on the NMR timescale. For example, a singlet resonance at δ 3.20 ppm (integrating as 18 protons) is observed in the 1H -NMR spectrum of **1** for the NMe_2 groups. We therefore conducted a variable temperature 1H -NMR study of **1** in toluene- d_8 from 298–193 K. The peaks in the NMR spectrum slowly broadened as the temperature was lowered, and the resonance at δ 3.20 ppm split into two broad peaks at δ 3.53 and 2.76 ppm (integrating in 2:1 ratio) at 213 K. The peak at δ 3.53 ppm split further into two broad peaks (at δ 3.59 and 3.48 ppm) at 203 K. Three fairly sharp resonances integrating in 1:1:1 ratio were observed at δ 3.60, 3.47, and 2.70 ppm at 193 K. The NMR data are consistent with a trigonal bipyramidal geometry about Ti with two equatorial NMe_2 ligands, one axial NMe_2 , and the bidentate salicylaldiminato ligand (L^1) coordinated at the remaining axial and equatorial sites. The molecular structure of $[L^1TiCl_3]$

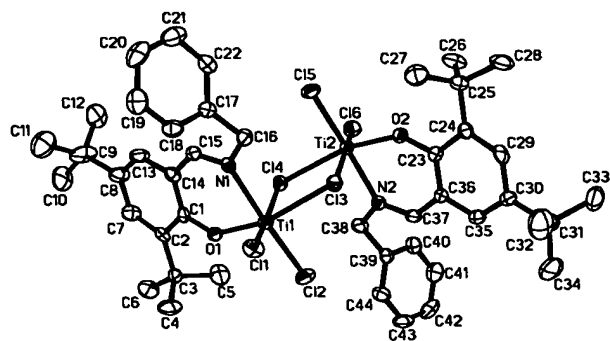


Fig. 2. An ORTEP diagram of the molecular structure of **3** showing 50% thermal ellipsoid probabilities.

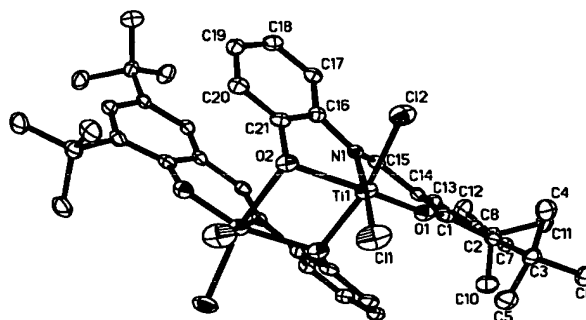


Fig. 3. An ORTEP diagram of the molecular structure of **4** showing 50% thermal ellipsoid probabilities.

(**2**), shown in Fig. 1, further supports the structural assignment made for $[L^1Ti(NMe_2)_3]$ (**1**). The geometry about the Ti center of **2** is trigonal bipyramidal, with one chloride and the imino nitrogen coordinated at the axial positions, and the remaining two chlorides and the aryloxy group coordinated at the equatorial sites. While the crystal of **2** utilized in the X-ray diffraction study was twinned and the poor crystal quality limits the accuracy of geometrical parameters, the connectivity is unambiguous and bond lengths and angles are within expected ranges 2h2i2j2k[11].

The six-coordinate complexes **3**, **4**, and **6** possess a distorted octahedral geometry about their Ti centers. The molecular structure of $[L^2TiCl_2(\mu-Cl)]_2$ (**3**), presented in Fig. 2, confirms the C_i symmetry of the molecule in solution (see 1H - and ^{13}C -NMR data) and metrical parameters for **3** (Table 2) are within the range observed for related six-coordinate, chloride-bridged Ti(IV) complexes 4a[11]. While mass spectral data allowed the formulation of $[L^3TiCl_2]_2$ (**4**) as a dimer ($m/z = 849$ for $[M^+ - Cl]$) and NMR data revealed a symmetric salicylaldiminato ligand environment, the molecular structure of **4** was unambiguously established

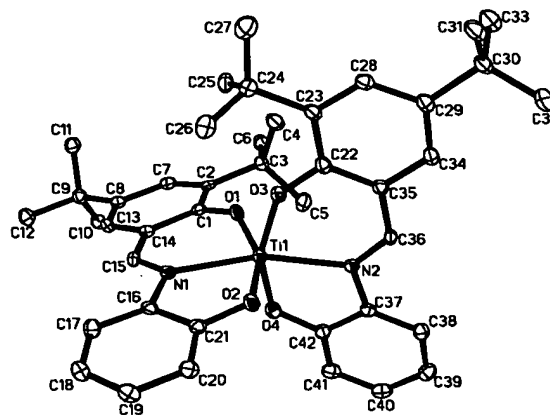


Fig. 4. An ORTEP diagram of the molecular structure of **6** showing 50% thermal ellipsoid probabilities.

Table 2
Selected bond distances (Å) and angles (°) for **3** and **4**

3		4 ^a	
<i>Bond distance</i>			
Ti(1) O(1)	1.783(4)	Ti(1) O(1)	1.7993(17)
Ti(1) N(1)	2.185(5)	Ti(1) O(2)	2.0573(16)
Ti(1) Cl(1)	2.2431(19)	Ti(1) O(2)#1	2.0934(16)
Ti(1) Cl(2)	2.2709(19)	Ti(1) N(1)	2.1697(19)
Ti(1) Cl(3)	2.4486(18)	Ti(1) Cl(2)	2.2708(7)
Ti(1) Cl(4)	2.5434(18)	Ti(1) Cl(1)	2.2780(7)
Ti(2) O(2)	1.791(4)	Ti(1) Ti(1)#1	3.2954(9)
Ti(2) N(2)	2.194(5)	N(1) C(15)	1.291(3)
Ti(2) Cl(6)	2.2518(19)	N(1) C(16)	1.423(3)
Ti(2) Cl(5)	2.2681(19)	O(1) C(1)	1.348(3)
Ti(2) Cl(4)	2.4379(18)	O(2) C(21)	1.382(3)
Ti(2) Cl(3)	2.5193(18)	O(2) Ti(1)#1	2.0934(16)
<i>Bond angles</i>			
O(1) Ti(1) N(1)	83.23(17)	O(1) Ti(1) O(2)	151.93(7)
O(1) Ti(1) Cl(1)	102.71(14)	O(1) Ti(1) O(2)#1	93.82(7)
N(1) Ti(1) Cl(1)	87.66(14)	O(2) Ti(1) O(2)#1	74.35(6)
O(1) Ti(1) Cl(2)	97.98(13)	O(1) Ti(1) N(1)	81.03(7)
N(1) Ti(1) Cl(2)	175.67(14)	O(2) Ti(1) N(1)	74.13(7)
Cl(1) Ti(1) Cl(2)	96.10(7)	O(2)#1 Ti(1) N(1)	91.67(6)
O(1) Ti(1) Cl(3)	159.28(14)	O(1) Ti(1) Cl(2)	98.77(6)
N(1) Ti(1) Cl(3)	86.27(13)	O(2) Ti(1) Cl(2)	93.37(5)
Cl(1) Ti(1) Cl(3)	94.62(7)	O(2)#1 Ti(1) Cl(2)	167.23(5)
Cl(2) Ti(1) Cl(3)	91.29(7)	N(1) Ti(1) Cl(2)	88.21(5)
O(1) Ti(1) Cl(4)	83.11(13)	O(1) Ti(1) Cl(1)	101.04(6)
N(1) Ti(1) Cl(4)	84.25(13)	O(2) Ti(1) Cl(1)	103.38(5)
Cl(1) Ti(1) Cl(4)	169.42(7)	O(2)#1 Ti(1) Cl(1)	86.16(5)
Cl(2) Ti(1) Cl(4)	91.76(6)	N(1) Ti(1) Cl(1)	177.09(6)
Cl(3) Ti(1) Cl(4)	78.09(6)	Cl(2) Ti(1) Cl(1)	93.48(3)

^a Symmetry transformations used to generate equivalent atoms: #1 $-x+2, 1-y, z$; #2 $y, 1-x, 1-z$; #3 $1-y, x, 1-z$.

Table 3
Selected bond distances (Å) and angles (°) for **6**

Bond distance	
< sp = 1/2 >	
Ti(1) O(3)	1.862(3)
Ti(1) O(1)	1.882(3)
Ti(1) O(4)	1.911(3)
Ti(1) O(2)	1.925(3)
Ti(1) N(2)	2.171(4)
Ti(1) N(1)	2.179(4)
Bond angles	
O(3) Ti(1) O(1)	88.44(13)
O(3) Ti(1) O(4)	155.76(14)
O(1) Ti(1) O(4)	90.34(13)
O(3) Ti(1) O(2)	93.04(13)
O(1) Ti(1) O(2)	156.97(13)
O(4) Ti(1) O(2)	97.42(13)
O(3) Ti(1) N(2)	81.29(14)
O(1) Ti(1) N(2)	112.58(14)
O(4) Ti(1) N(2)	76.86(14)
O(2) Ti(1) N(2)	90.34(13)
O(3) Ti(1) N(1)	110.24(13)
O(1) Ti(1) N(1)	81.29(13)
O(4) Ti(1) N(1)	93.47(13)
O(2) Ti(1) N(1)	76.64(13)
N(2) Ti(1) N(1)	162.78(14)

by X-ray crystallography (Fig. 3) [12]. The Ti centers of **4** are surprisingly bridged by the oxygen atom of the sterically less hindered aryloxy moiety of each salicylaldiminato ligand. Equally interesting, the salicylaldiminato ligands are oriented such that the phenyl ring of the less substituted aryloxy moiety of one ligand is face to face with the phenyl ring of the more substituted aryloxy moiety of the second salicylaldiminato ligand. The distance from plane of the ring labeled C16–C21 to the centroid of the di-*tert*-butyl-substituted ring facing it is 3.293(3) Å, consistent with a π – π stacking interaction [13]. Bond distances and angles of **4** (Table 2) are within the expected ranges although the Ti to terminal aryloxy oxygen bond is somewhat short [Ti(1)–O(1) = 1.7993(17) Å]. This probably reflects increased donation to Ti by this aryloxy oxygen due to its orientation trans to the bridging aryloxy oxygen [Ti(1)–O(2) is much longer at 2.0573(16) Å]. That the bridging aryloxy is weakly-bound is evidenced by the quantitative formation of monomeric [L³TiCl₂(THF)] (**5**) when **4** was dissolved in THF and stirred for 16 h (vide infra). The molecular structure of [(L³)₂Ti] (**6**) revealed that the tridentate salicylaldiminato ligands are coordinated in a meridional fashion and orthogonal to one another (Fig. 4; selected bond distances and angles are presented in Table 3). In addition, the structure provides support for the molecule being C₂-symmetric in solution, as deduced on the basis of ¹H- and ¹³C-NMR data (see Section 2).

With methylalumoxane (MAO) (900 molar equivalents) as co-catalyst, **1** and **2** showed ethylene polymerization activities at 25 °C (Table 4) comparable to those reported for several noncyclopentadienyl-based Ti and Zr systems [14]. Thus, **1** and **2** are much less effective catalyst precursors than Cp₂ZrCl₂, which was 28 times more active under identical polymerization conditions, or Group 4 metal complexes based on chelating di(amido)- 2c2d2f], amine-bis(phenolato)- 2n, or bis(salicylaldiminato) 2h2i2j2k ligands. The modest ethylene polymerization activities may be due to several factors, including a low equilibrium concentration of the putative active cationic species, [LTiMc₂]⁺ (L = L¹ or L²). We presume that **2** is a somewhat more effective catalyst precursor because the active catalyst species can be more easily generated by chloride substitution. The fact that mono(salicylaldiminato) complexes **1** and **2** are much less effective catalyst precursors than bis(salicylaldiminato)Ti(IV) complexes probably reflects a decreased stability of the active cationic species, which will be more electron deficient and stabilized to a less degree by sterics than the active species generated from bis(salicylaldiminato)Ti(IV) complexes. The stabilization of the cationic species would lead to an increased concentration of the active catalyst and hence to higher activity 2n[14].

Table 4
Ethylene polymerization at 25 °C

Catalyst	mmol of catalyst	Molar excess of MAO	Time (min)	Activity (kg molcat ⁻¹ h ⁻¹)
1 ^a	0.028	900	10	14
2 ^a	0.027	900	10	30
Cp ₂ ZrCl ₂ ^a	0.025	900	10	835
2 ^b	0.025	900	10	9

^a 1 atm C₂H₄ pressure bubbled through toluene mixture of catalyst/MAO, which had been stirred under N₂ atmosphere for 30 min.

^b 1 atm C₂H₄ pressure bubbled through toluene mixture of catalyst/MAO immediately after mixing.

4. Conclusions

Monomeric mono(salicylaldiminato)Ti(IV) complexes can be prepared through appropriate choice of the salicylaldiminato ligand. With MAO as co-catalyst, Ti(IV) tris(dimethylamide) and -trichloride complexes **1** and **2** showed modest activities in ethylene polymerization. The complexes are much less effective catalyst precursors than previously reported bis(salicylaldiminato) complexes of the Group 4 metals. The design and study of related complexes are currently underway in our laboratory, with the aim of developing highly active olefin polymerization catalysts.

5. Supplementary material

Crystallographic data for the structural analysis have been deposited with the Cambridge Crystallographic Data Centre, CCDC nos. 204405–204408 for **2**, **3**, **4**, and **6**, respectively. Copies of the information may be obtained free of charge from The Director, CCDC, 12 Union Road, Cambridge CB2 1EZ, UK (Fax: +44-1223-336033; email: deposit@ccdc.cam.ac.uk or www: http://www.ccdc.cam.ac.uk).

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